

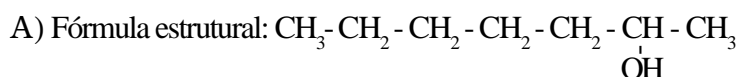
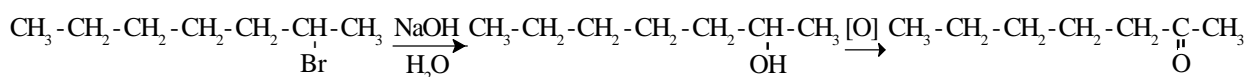
CONTEÚDO

3

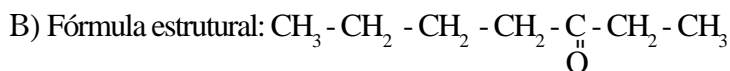
A) $[\text{OH}^-] = 10^{-5} \text{ mol.L}^{-1}$
 $\text{pOH} = -\log [\text{OH}^-] = -\log 10^{-5} = 5$
 $\text{pH} + \text{pOH} = 14$
 $\text{pH} = 14 - 5 = 9$
 $\text{pH} = 9$

B) A acidez do limão tende a neutralizar o meio básico, diminuindo a concentração de íons OH^- e deslocando o equilíbrio para o lado direito. Isso provoca a diminuição da concentração da metilamina, que é a substância responsável pelo odor característico do peixe.

4

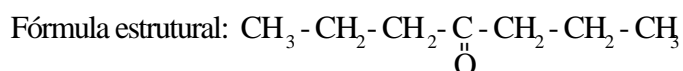


Nomenclatura: 2-heptanol



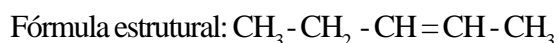
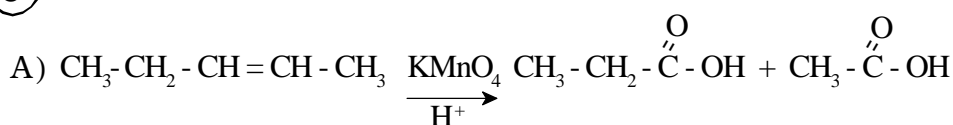
Nomenclatura: 3-heptanona

ou

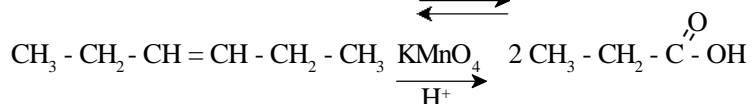
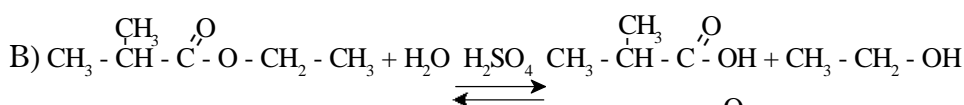


Nomenclatura: 4-heptanona

5



Número de carbonos secundários: 3



Fórmula molecular: $\text{C}_3\text{H}_6\text{O}_2$

Nomenclatura: propanóico